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## Periodic Trends Answers

**periodic trends worksheet answers** - periodic(trends(worksheet(((answer key 1. circle)the)element)withthe)largest)atomic(radius)andput)a)square)aroundthe)element)withthe)smallest)atomic(radius:))) cu) k) ni) br))) largest - k smallest - br) explainwhy)you)madethesechoices.) atomic radius decreases as you go left to right across a period. potassium is in the far **periodic trends: electronegativity answers. name** - periodic trends: electronegativity answers. name \_\_\_\_ 1. what is the electronegativity of an element? it is the tendency of a bonded atom to accept or to donate an electron. 2. is the "electronegativity" the same thing as the "electron affinity"? no, the electronegativity relates to bonded atoms and their tendency to attract ... **periodic trends worksheet - sd308** - answers for comparing atomic sizes here are answers for the questions above. a. li, c, f all are in the same period and thus have the same number of energy levels. therefore, the important factor is the nuclear charge. li is the largest because it has the smallest nuclear charge and pulls the ... microsoft word - periodic trends worksheet author: **answer key trends on the periodic table - rocklin.k12** - answer key trends on the periodic table part 1: identifying chemical properties/characteristics of the elements for each of the following elements, draw a caricature or cartoon above the description to show its **periodic trends quiz - howechem** - periodic trends quiz 1. from which of the following is it easiest to remove an electron? a. mg b. na c. al d. s answer: na --it's the biggest 2. which of the following influenced your answer to number one the most? a. nuclear charge (number of protons) b. valence electrons c. inner shell electrons d. shielding effect **6.3 periodic trends - henry county schools / overview** - section 6.3 periodic trends 171 group trends in atomic size in the figure 6.14 graph, atomic radius is plotted versus atomic number. look at the data for the alkali metals and noble gases. the atomic radius within these groups increases as the atomic number increases. this increase is an example of a trend. **name worksheet: periodic trends period - kmacgill** - name\_\_\_\_ worksheet: periodic trends period\_\_\_\_ 1. which statement best describes group 2 elements as they are considered in order from top to bottom of the periodic table? (a) the number of principal energy levels increases, and the number of valence electrons increases. **periodic trends multiple choice review psi chemistry name** - njctl periodic trends chemistry periodic trends multiple choice review psi chemistry name\_\_\_\_ atomic size 1) elements z and x are compared. element z is larger than element x. based on this you could say: a) element z is further to the left side of the periodic table **periodic trends in reactivity lab - science curriculum** - periodic trends in reactivity list the four metals from most reactive to least reactive. use data from your lab to support your answer. in general, is there a relationship between the locations of metals on the

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