## **Periodic Trends Unit Test Answers**

chemistry 11 unit review: atomic theory & periodic trends - chemistry 11 unit review: atomic theory & periodic trends atoms atoms have protons and neutrons located in the nucleus of the atom. electrons orbit around the nucleus in well-defined paths, protons have positive charge, electrons have negative charge, and neutrons are neutral, greek philosophers (democritus and aristotle 400-500 bc) periodic trends quiz **howechem** - periodic trends quiz 1. from which of the following is it easiest to remove an electron? a. mg b. na c. al d. s answer: na --it's the biggest 2. which of the following influenced your answer to number one the most? a. nuclear charge (number of protons) b. valence electrons c. inner shell electrons d. shielding effect periodic trends worksheet - sd308 - worksheet: periodic trends 1. atomic radius for each of the following sets of atoms, rank the atoms from smallest to largest atomic radius. a. li, c, f b. li, na, k c. ge, p, o d. c, n, al e. al, cl, ga 2. ionic radius for each of the following sets of ions, rank them from smallest to largest ionic unit 2 the periodic table - unit 2 - the periodic table this unit will be both quick and easy, serving as a bridge between two very challenging units. objectives: at the end of this unit, students will: understand the placement of each of the elements on the periodic table and be able to predict the location of an "unknown" element based on its chemical and physical unit 7 test: periodic table - weebly - unit 7 test: periodic table 1. this law states that when elements are arranged in order of increasing atomic number their physical and chemical properties show a periodic pattern. a. mendeleev's law c. periodic law b. mosely's law d. family law 2. cations have a a, positive charge b, negative charge c, neutral charge d, unit 3 the periodic table, electron configuration ... - unit 3 -the periodic table, electron configuration, &periodic trends chapters 4 & 5. creation of the periodic table, mendeleev's table dmitry mendeleev he created the first periodic table based on the properties of the elements. chemistry 11 unit review: atomic theory & periodic trends - chemistry 11 unit review: atomic theory & periodic trends atoms atoms have protons and neutrons located in the nucleus of the atom. electrons orbit around the nucleus in well-defined paths. protons have positive charge, electrons have negative charge, and neutrons are neutral. periodic table and trends test review key - mr. b. gillis ... - periodic table and trends test review key describe the common characteristics and uses of the following families: 1. alkali metals (# of valence electrons = 1) • most reactive metals that do not occur freely in nature • slivery-white and softer than most other metals (to the point that they can be cut easily with a knife) the periodic table and atomic structure test - the periodic table and atomic structure test multiple choice identify the choice that best completes the statement or answers the question. 31) the smallest particle into which an element can be divided and still have the properties of that element is a.p. chemistry practice test - ch. 7, atomic structure and ... - a.p. chemistry practice test - ch. 7, atomic structure and periodicity multiple choice, choose the one alternative that best completes the statement or answers the question. 1) ham radio operators often broadcast on the 6-meter band, the frequency of this electromagnetic mhz. periodic trends worksheet answers - periodic(trends(worksheet(((answer key 1. circle)the)element)withthe)largest)atomic)radius)andput)a)square)aroundthe)element)withthe)smallest) atomic)radius:))) cu) k) ni) br))) largest - k smallest - br) explainwhy)you)madethesechoices.) atomic radius decreases as you go left to right across a period. potassium is in the far periodic trends multiple choice review psi chemistry name - njctl periodic trends chemistry periodic trends multiple choice review psi atomic size 1) elements z and x are compared. element z is larger than element x. based chemistry name on this you could say: a) element z is further to the left side of the periodic table practice test for unit 4 chemistry i 2012 - 1&1 internet - summarize the same trends as in #2 as you move across a period in the periodic table. explain why these trends occur. 4. rank the following atoms and ions in size from least to greatest. explain your answer. a. se se-2 br b. se-2 br i c. al al+3 c practice test for unit 4 chemistry i 2012 1 periodic practicetest2012-13.odt unit 3 atomic structure and periodic table - unit 3 note packet and goals unit 3 - atomic structure and periodic table unit goals- as you work through this unit, you should be able to: 1. describe previous atomic theories and compare to our modern understanding of the atom (4.1) 2. distinguish among protons, electrons, and neutrons in terms of mass and charge. chemistry--chapter 5: atomic structure and the periodic table - chemistry--unit 1: atomic structure and the periodic table test review vocab o atom o atomic mass o atomic mass unit o atomic number o electron o isotopes o mass number o metals o metalloids o neutrons o nonmetals o periodic law o proton know 1. dalton's atomic theory and what part is now known to be different 2.

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